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DISCUSSION.

Dr. T. M. Lowry said that the abnormal electrode-potentials of silver in contact with silver acetate found a parallel in the electrochemical equivalent for silver deposited from this solution, which was abnormally high, rising to 1·12 mg. per coulomb. The deposit did not consist of pure silver, however, but was backed by a yellowish film of some other material; possibly, some similar effect was responsible for the curious observations recorded by the author in the case of this salt.

Dr. E. Feilmann said it was almost impossible to obtain hydrogen free from zinc if made from that metal. It was therefore quite likely that the

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hydrogen used by the author and blown into the solution contained minute traces of zinc, and this might account for the variations in potential observed. Of course this explanation would not apply to the cases of nitrogen or carbon dioxide.

Dr. N. T. M. Wilsmore said that if the effect were due to the presence of zinc, it should take place in the case of the Ag-$AgNO_3$ electrode. Besides, if due to the reduction of silver, it would produce an opposite effect to that observed.

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